



CANDIDATE
NAME

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CENTRE
NUMBER

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NUMBER

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0620/31

May/June 2021

1 hour 15 minutes

You must answer on the question paper.

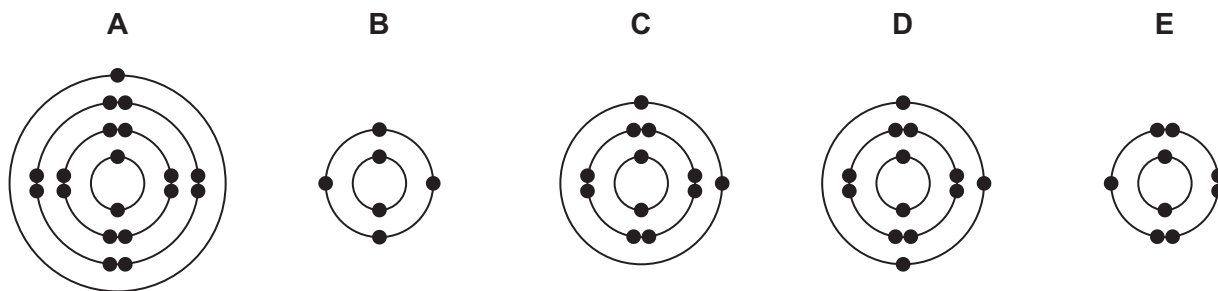
No additional materials are needed.

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [].
- The Periodic Table is printed in the question paper.

This document has **16** pages. Any blank pages are indicated.

- 1 (a) The electronic structures of five atoms, **A**, **B**, **C**, **D** and **E**, are shown.



Answer the following questions about these electronic structures.
Each electronic structure may be used once, more than once or not at all.

State which electronic structure, **A**, **B**, **C**, **D** or **E**, represents:

- (i) an atom in Group II of the Periodic Table

..... [1]

- (ii) an atom with a proton number of 13

..... [1]

- (iii) an atom that forms a stable ion with a single negative charge

..... [1]

- (iv) an atom of a non-metal that forms a giant covalent structure

..... [1]

- (v) an atom of a metal used in food containers.

..... [1]

- (b) Complete the table to show the number of electrons, neutrons and protons in the vanadium atom and calcium ion shown.

	number of electrons	number of neutrons	number of protons
$^{51}_{23}\text{V}$	23		
$^{48}_{20}\text{Ca}^{2+}$		28	

[3]

[Total: 8]

- 2 The table shows the masses of some of the ions in 1000 cm^3 of fruit juice.

name of ion	formula of ion	mass of ion in 1000 cm^3 of fruit juice / mg
	NH_4^+	43
calcium	Ca^{2+}	79
chloride	Cl^-	135
lithium	Li^+	1
magnesium	Mg^{2+}	80
nitrate	NO_3^-	35
phosphate	PO_4^{3-}	120
potassium	K^+	575
sodium	Na^+	120
	SO_4^{2-}	105

- (a) Answer these questions using only the information in the table.

- (i) State which negative ion has the highest mass in 1000 cm^3 of fruit juice.

..... [1]

- (ii) Give the formulae of the ions in ammonium sulfate.

..... and [1]

- (iii) Calculate the mass of sodium ions in 200 cm^3 of fruit juice.

mass = mg [1]

- (b) Describe a test for lithium ions.

test

observations

[2]

- (c) Ions of the element potassium, K, are present in most fertilisers.

State the names of two **other** elements that are in most fertilisers.

1

2

[2]

- (d) Orange juice is acidic.

Draw a circle around the pH of orange juice.

pH 4

pH 7

pH 10

pH 13

[1]

- (e) Some soils are acidic.

Give the names of **two** compounds that are used to make soils less acidic.

1

2 [2]

- (f) Hydrogen chloride is an acidic gas produced when concentrated hydrochloric acid evaporates.

- (i) Describe the arrangement and separation of the molecules in hydrogen chloride gas.

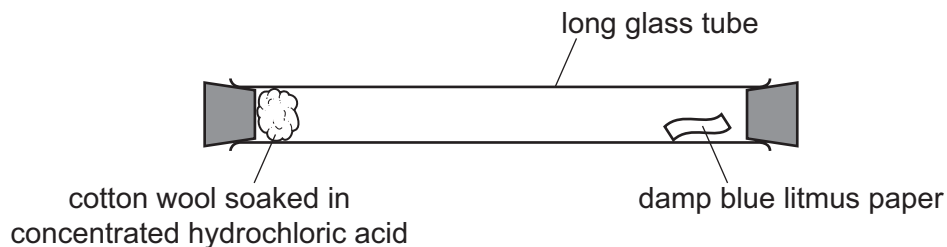
arrangement

.....

separation

..... [2]

- (ii) A long glass tube is set up as shown.



At first, the blue litmus paper does not turn red.

After a short time, the litmus paper turns red.

Explain these observations using the kinetic particle model.

.....

.....

.....

.....

..... [3]

[Total: 15]

3 The table shows some properties of four Group I elements.

element	melting point /°C	boiling point /°C	atomic radius /nm
sodium	98	883	0.191
potassium	63	760	
rubidium	39		0.250
caesium	29	671	0.272

(a) (i) Complete the table by predicting:

- the boiling point of rubidium
- the atomic radius of potassium.

[2]

(ii) Describe the trend in the melting point of the Group I elements down the group.

..... [1]

(iii) Deduce the physical state of potassium at 60 °C.
Explain your answer.

.....

 [2]

(b) Caesium is a radioactive element with a proton number of 55.

(i) Define proton number.

..... [1]

(ii) State **one** industrial use of radioactive isotopes.

..... [1]

(c) Sodium hydride, NaH, reacts with iron(III) oxide.

(i) Balance the equation for this reaction.

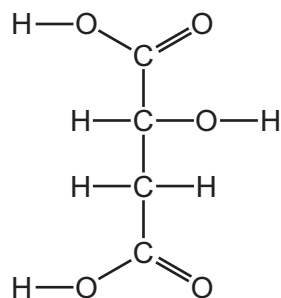


(ii) Explain how this equation shows that iron(III) oxide is reduced.

..... [1]

[Total: 10]

4 The structure of malic acid is shown.



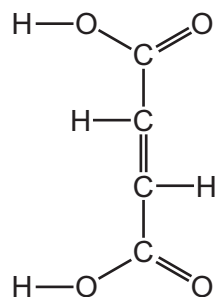
(a) (i) On the structure draw a circle around the alcohol functional group. [1]

(ii) Deduce the formula of malic acid to show the number of carbon, hydrogen and oxygen atoms.

..... [1]

(b) When malic acid is heated it forms compound **F**.

The structure of compound **F** is shown.



Explain why compound **F** is described as unsaturated.

..... [1]

(c) Compound **F** can form polymers.

(i) State the meaning of the term *polymer*.

..... [2]

(ii) State the name of the polymer formed when ethene is polymerised.

..... [1]

(d) Ethanoic acid is a carboxylic acid.

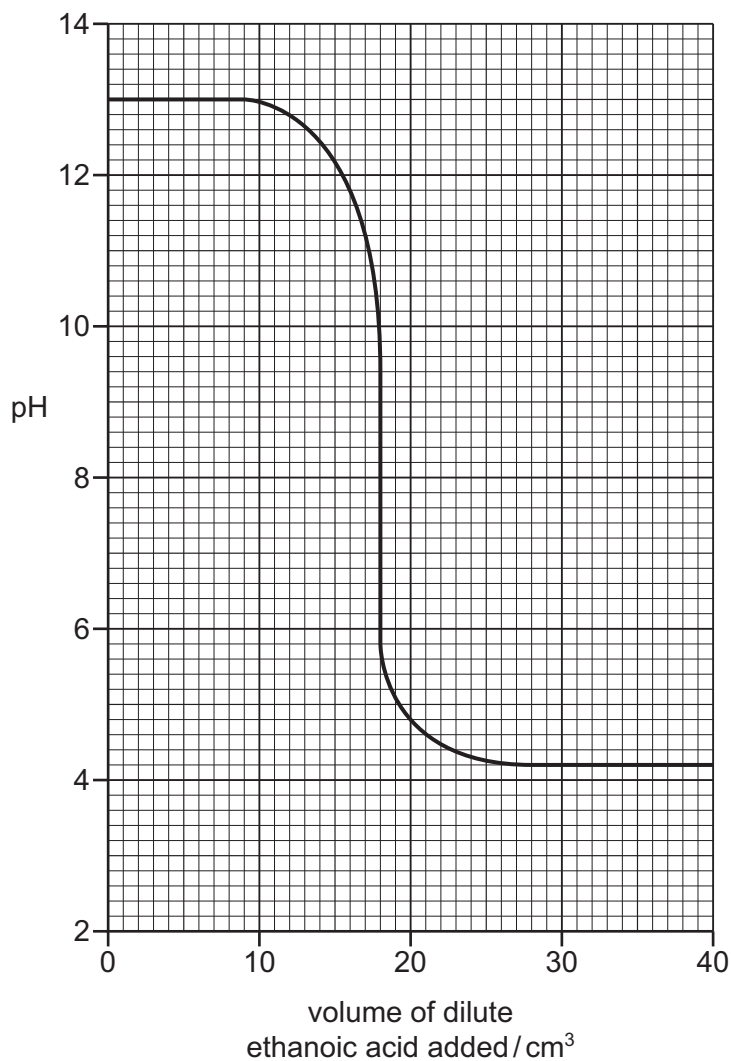
Describe the observations made when dilute ethanoic acid reacts with:

magnesium

litmus solution.

[2]

(e) The graph shows how the pH changes when dilute ethanoic acid is added slowly to aqueous sodium hydroxide.



(i) Deduce the pH of the aqueous sodium hydroxide before the addition of dilute ethanoic acid.

pH = [1]

(ii) Deduce the volume of dilute ethanoic acid added when the pH is neutral.

..... cm³ [1]

[Total: 10]

5 (a) Calcium oxide is made by the thermal decomposition of calcium carbonate.

(i) State the meaning of the term *thermal decomposition*.

.....
..... [2]

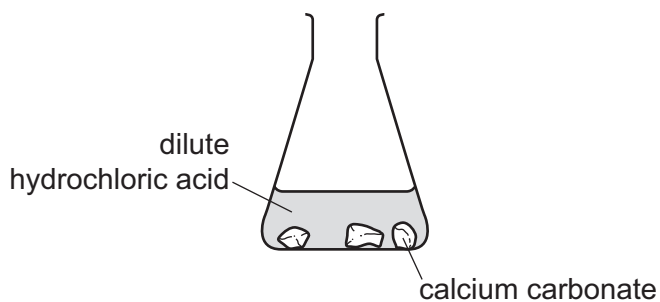
(ii) Describe a test for calcium ions.

test
observations [2]

(b) Carbon dioxide is produced when dilute hydrochloric acid reacts with calcium carbonate.



(i) Complete the diagram to show how to measure the volume of carbon dioxide produced during this reaction.



[2]

- (ii) Describe the effect of each of the following on the rate of reaction of dilute hydrochloric acid with calcium carbonate.

- The concentration of hydrochloric acid is decreased.

All other conditions stay the same.

.....

- The temperature is increased.

All other conditions stay the same.

.....

[2]

- (c) Carbon dioxide is also formed when the hydrocarbon C_3H_8 is completely combusted.

- (i) State the meaning of the term *hydrocarbon*.

.....

..... [2]

- (ii) The hydrocarbon C_3H_8 is called propane.

Name the homologous series that propane belongs to.

..... [1]

- (iii) Name two substances formed by the incomplete combustion of propane.

..... and [2]

[Total: 13]

6 This question is about water.

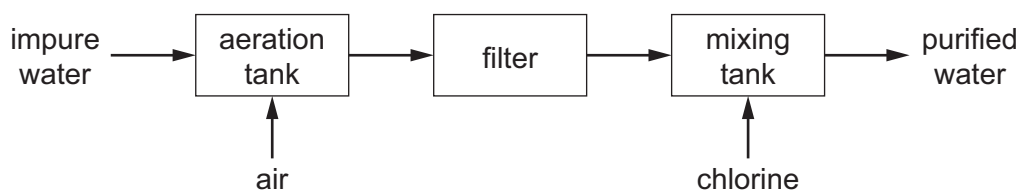
(a) The water in rivers often contains pollutants such as acids.

Describe how universal indicator paper can be used to determine the pH value of the water.

.....

 [2]

(b) The diagram shows some of the stages in water treatment.



(i) Air is blown through the aeration tank.

Name the **two** gases that make up most of the air.

..... and [2]

(ii) After aeration, the water still contains large insoluble particles.

The filter is made up of fine sand and stones.

Explain how the filter helps purify the water.

.....

 [2]

(iii) Explain why chlorine is used in water treatment.

..... [1]

(c) Anhydrous cobalt(II) chloride is used to test for water.

State the colour change in this test.

from to [2]

[Total: 9]

7 (a) Molten lead(II) bromide is electrolysed using carbon electrodes.

(i) State the products of this electrolysis at:

the negative electrode

the positive electrode.

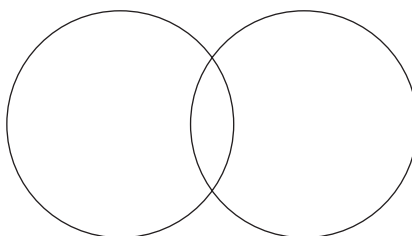
[2]

(ii) State the name of **another** substance which can be used as an inert electrode.

..... [1]

(b) When aqueous magnesium chloride is electrolysed using carbon electrodes, hydrogen gas is produced at the negative electrode.

Complete the dot-and-cross diagram to show the electron arrangement in one molecule of hydrogen.



[1]

[Total: 4]

8 This question is about elements in the Periodic Table.

(a) The table shows some properties of five elements, **P**, **Q**, **R**, **S** and **T**.

element	melting point /°C	density in g/cm ³	electrical conductivity of the solid	atomic radius /nm
P	1535	7.86	very good	0.125
Q	−7	3.12	does not conduct	0.114
R	1495	8.90	very good	0.126
S	−157	0.0035	does not conduct	0.110
T	839	1.54	very good	0.174

Use only the elements shown in the table to answer this question.

State which two of the elements, **P**, **Q**, **R**, **S** and **T**, are covalent molecules.
Give **two** reasons for your answer.

elements and

reason 1

reason 2

[3]

(b) Element **T** is on the left-hand side of the Periodic Table.
Suggest whether its oxide is acidic or basic.

Give a reason for your answer.

.....

..... [1]

(c) Krypton is an element in Group VIII of the Periodic Table.

Explain, using ideas about electronic structure, why krypton is unreactive.

.....

..... [1]

- (d) Sodium is an element in Group I of the Periodic Table. Iron is a transition element.

Iron has a higher melting point and higher boiling point than sodium.

Give **two** other ways in which the properties of transition elements differ from the properties of Group I elements.

1

2

[2]

- (e) The table compares the reactivity of four metals with dilute hydrochloric acid.

metal	reaction with dilute hydrochloric acid
calcium	reacts very rapidly
copper	no reaction
iron	reacts rapidly
nickel	reacts slowly

Put the four metals in order of their reactivity.

Put the least reactive metal first.

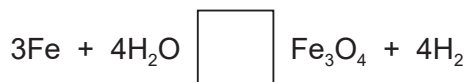
least reactive $\longrightarrow$ most reactive

--	--	--	--

[2]

- (f) Hot iron reacts with steam. The reaction is reversible.

Complete the equation by writing the symbol for a reversible reaction in the box.



[1]

- (g) Steel is an alloy of iron.

State the meaning of the term *alloy*.

..... [1]

[Total: 11]

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The Periodic Table of Elements

Group																			
I	II											III	IV	V	VI	VII	VIII		
		1Hhydrogen1																	
		Key atomic number atomic symbol name relative atomic mass																	
3Li lithium 7	4Be beryllium 9													5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	10Ne neon 20
11Na sodium 23	12Mg magnesium 24													13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84		
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131		
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids		72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —		
87Fr francium —	88Ra radium —	89–103 actinoids		104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	114Fl flerovium —	116Lv livermorium —					

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).